

Mixed Stoichiometry Practice

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Play this game to review Chemical Reactions. Using the following equation: $\text{Fe}_2\text{O}_3(\text{s}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{Fe}(\text{s}) + 3\text{H}_2\text{O}(\text{l})$ How many moles of Fe can be made from 6 moles H₂? (This is a one step conversion using mole ratio)

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Stoichiometry Stoichiometry is the practice of predicting the amount

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of product or reactant in a chemical equation based on a known amount of one of the other products or reactants. The ability to do these calculations is the culmination of all of the basic skills learned throughout the first semester.

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Stoichiometry: Mixed Problems (KEY) 1) $N_2 + 3H_2 \rightarrow 2NH_3$ What volume of NH_3 at STP is produced if 25.0 of N_2 is reacted with an excess of H_2 ?
3 3 3 2 3 2 2 2 40.0L NH_3 1mol NH_3 22.4L NH_3 1mol N_2 2mol NH_3 28.0g N_2 25.0g
 N_2 1mol N_2 $\times \times \times =$ 2) $2KClO_3 \rightarrow 2KCl + 3O_2$ If 5.0g of $KClO_3$ is decomposed, what volume of O_2 is produced at STP? 2

~~Stoichiometry: Mixed Problems (KEY)~~

Mixed Stoichiometry Problems . 1. $2H_2 + O_2 \rightarrow 2H_2O$. a). How many moles of H_2 would be required to produce 5.0 moles of water? 5.0 moles water. b). What mass of H_2O is formed when H_2 reacts with 384 g of O_2 ? 432g H_2O . 2. $H_2SO_4 + 2NaOH \rightarrow Na_2SO_4 + 2H_2O$. a). Balance this equation. Look above. b).

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Learn how to use mole ratios derived from balanced chemical equations to calculate amounts of substances consumed and produced in chemical

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reactions.

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Stoichiometry / , s t o i k i ' o m i t r i / is the calculation of
reactants and products in chemical reactions.. Stoichiometry is
founded on the law of conservation of mass where the total mass of the
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